

SPECIFICATION

WATER-BASED LUBRICANTS CONTAINING SULFUR AS A COORDINATING ATOM AND USES THEREOF

INSA, >

5 Technical Field

10 The present invention relates to aqueous lubricants that are coated onto either the surface of a metal material or the molding surface of a metal mold, or both, to form a lubricating film on the metal surface, so that friction between the material and the mold is reduced, thereby enabling satisfactory plastic working of the metal. The present invention further relates to methods of using the aqueous lubricant and to chemical substances that are particularly suited for production of the aqueous lubricant.

15 Background Art

20 When a metal material is subjected to plastic working, such as forging, extrusion, drawing, rolling or pressing, a lubricant film must be disposed on the surface of the metal material and/or the molding surface of the metal mold to reduce friction between the material and the mold and to prevent the generation of drag, seizures and the like.

25 In most cases, the metal is immersed in a zinc phosphate solution to produce a zinc phosphate or iron phosphate chemical film (hereafter referred to as "phosphate film") on the surface,

and then further immersed in a sodium soap or the like to produce a metallic soap layer on the surface. This process creates a surface protective layer of satisfactory quality, and allows heavy working (meaning working of material surfaces with large elongation rates). However, this process requires washing treatments with cold water, hot water or acid, as well as different types of equipment. Such treatments also require long periods of time, thus lengthening the lead time. Usually, 30 minutes or longer is necessary to complete a series of treatments. It is also necessary to accomplish a single surface treatment of the entire portion of the metal that is subjected to plastic working in a subsequent step during the treatment time, and therefore this process is not suitable for production of small, sundry products.

In order to overcome these problems, the use of working oils has been proposed. For example, Japanese Laid-Open Patent Publication No. 7-118682 teaches a working oil prepared by dispersing in a mineral oil a zinc or molybdenum salt, such as zinc dithiophosphate, molybdenum dithiocarbamate or the like, which has been rendered lipophilic by the introduction of a higher alkyl group. This working oil solves most of the aforementioned problems, but because the main component is oil, it creates problems, such as contamination due to adhesion of the oil to surrounding machines and generation of oil mist. In other words, contamination of the working environment is unavoidable. A

number of other problems also remain, such as the need to degrease the material surface after plastic working.

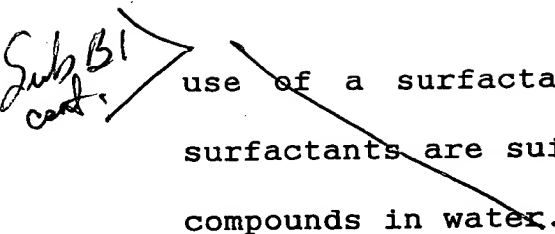
Disclosure of the Invention

5 The present invention overcomes the problems mentioned above, and allows formation of lubricating films on surfaces by application of aqueous lubricants containing no oil. Because no oil is used, the problems of working environment contamination and the need for subsequent degreasing treatment are solved. Because a simple application is sufficient, a solution is also provided to the problems of large equipment requirements and the need to accomplish simultaneous treatment of large volumes of materials.

10
15
20
25

Sub B1

The invention according to claim 1 relates to the aqueous lubricant itself, in which the aqueous lubricant comprises a metal chelate compound suspended or dispersed in water. The metal chelate compound described herein comprises a polydentate or multidentate chelate ligand in which at least one of the coordinating atoms is sulfur, coordinated to a coordination site of at least one metal species selected from among zinc, manganese, iron, molybdenum, tin and antimony. For the purposes of this explanation, the term "suspension" is intended to mean a metal chelate compound distributed in water, for example, by continuous stirring. The term "dispersion" is intended to mean a metal chelate compound distributed in water without precipitation, by

Sub B1
cont.  use of a surfactant or the like. Anionic and non-ionic surfactants are suitable for dispersion of the metal chelate compounds in water.

10
15
The metal may have multiple coordination sites and the chelate ligand may be coordinated to all the multiple coordination sites. Alternatively, the chelate ligand may be coordinated to only some of the multiple coordination sites, thereby allowing coordination of a species other than the chelate ligand to the remaining coordination sites. The lubricant can be produced by dispersing the prepared metal chelate compound in water, or it may be produced by adding the chelate ligand to an aqueous solution containing a metal salt. The metal species may be any of one or more desired species, and a greater number of species will expand the range of workable conditions and workable metals.

20
25
The aqueous lubricant is applied onto the surface of a metal material and/or the molding surface of a metal mold to form an effective lubricating film on the metal surface. Because the lubricating film has sulfur as a coordinating atom, extreme pressure produces sulfur radicals through decomposition by friction or tribo-chemical reactions. The sulfur radicals are highly reactive, and react rapidly with the metal surface to produce metal sulfides, which have a lubricating effect. The sulfur radicals also react with the metal ions (one or more selected from among zinc, manganese, iron, molybdenum, tin and

antimony) produced by decomposition of the metal chelate compound, also producing metal sulfides having a lubricating effect. The aqueous lubricant thus exhibits a satisfactory lubricating effect.

5
Sub B2
The invention according to claim 2 relates to chemical substances particularly suited for production of the aqueous lubricant. The chemical substances are multi-ligand metal chelate compounds, in which a polydentate or multidentate chelate ligand having sulfur as at least one of the coordinating atoms coordinates by partially filling the multiple coordination sites of the one or more metal species selected from among zinc, manganese, iron, molybdenum, tin and antimony, whereas ligands that do not have sulfur as a coordinating atom coordinate to the remaining coordination sites. That is, the chelate ligand is characterized as having sulfur as a coordinating atom that does not fill all the coordination sites of the metals, so that it is not coordinated to some of the coordination sites. When the multi-ligand metal chelate compound is used as an aqueous lubricant suspended or dispersed in water, a very satisfactory
10
15
20 lubricating film is produced.

Sub B3
The invention according to claim 3 also relates to a chemical substance particularly suited for production of the aqueous lubricant. The chemical substance is characterized in that a chelate ligand having sulfur as a coordinating atom is
25 coordinated to some of the multiple coordination sites of the

Sub B3
cont. metal, whereas a hydroxide ion, condensed phosphate, polycarboxylic high molecular activator and/or polyoxycarboxylic acid are coordinated to the remaining coordination sites.

5 In the multi-ligand metal chelate compound, a chelate ligand having sulfur as a coordinating atom is strongly coordinated with the metal, and a hydroxide ion, condensed phosphate, polycarboxylic high molecular activator and/or polyoxycarboxylic acid are weakly coordinated with the metal via oxygen anions. When the multi-ligand metal chelate compound is used as an aqueous lubricant suspended or dispersed in water, a very satisfactory lubricating film is produced.

Sub B4
10 The aqueous lubricant according to claim 5 has a soluble condensed phosphate salt, a soluble polycarboxylic high molecular activator and/or a soluble polyoxycarboxylic acid salt added to the aforementioned aqueous lubricant. Addition of these adjuvants improves the performance of the lubricating film. A soluble condensed phosphate salt will associate with the surfactant present in the system, thus increasing the
15 dispersability of the metal chelate compound that forms hydrophobic fine particles. A soluble polycarboxylic high molecular activator or soluble polyoxycarboxylic acid salt will increase the adhesion of the lubricating film to the metal surface. Using an aqueous lubricant containing such adjuvants will allow
20 more intense heavy working.
25

Sub B5

The invention according to claim 6 relates to a process of forming a lubricating film on a phosphate film using an aqueous solution, if the phosphate film had already been formed on a metal surface. In this process, a metal material on which the phosphate film has already been formed is immersed in an aqueous solution of a multidentate or polydentate chelate ligand having sulfur as at least one of the coordinating atoms, so that the chelate ligand reacts with the zinc ion and/or iron ion in the phosphate film to produce a crystalline multi-ligand metal chelate compound on the phosphate film. This process takes advantage of both the lubricating effect of the phosphate film and the lubricating action of the metal chelate compound, in which sulfur is a coordinating atom chelated to the zinc ion and/or iron ion.

10
11
12
13
14
15
16
17
18
19
20
21
22
23
24
25

Sub B6

The invention according to claim 7 also relates to a process of forming a lubricating film on a phosphate film. In this process, the phosphate film is formed on a metal material and the metal material is then immersed in an aqueous lubricant according to claim 4 or 5. A ligand, which is not a ligand having sulfur as a coordinating atom, reacts with the zinc ion and/or iron ion in the phosphate film to produce a crystalline polynuclear metal chelate compound on the phosphate film. This process takes advantage of both the lubricating effect of the phosphate film and the lubricating action of the metal chelate compound, in which sulfur is a coordinating atom chelated to a metal.

Sub B1

The invention according to claim 8 relates to a method of using the aqueous lubricant, in which prior to plastic working of the metal material, an aqueous lubricant according to claim 1, 4 or 5 is applied onto either or both surfaces of the metal material and the molding surface of the metal mold to form lubricating films on those surfaces, thus allowing plastic working of the metal material with a lubricating film formed on the surface.

The invention will be more fully appreciated in light of the explanation that follows.

Best Mode for Practicing the Invention

First, an embodiment of a metal chelate compound will be explained in which at least one metal species selected from among zinc, manganese, iron, molybdenum, tin and antimony is chelated and at least one of the coordinating atoms is sulfur. Chemical structures 1-28 below are examples of such metal chelate compounds, and the chelate ligands in chemical structures 1-28 (the compounds adjacent to M in the structures) react with the above-mentioned metal ions in aqueous solution or in water or organic solvents (alcohols, ketones and dioxane) to produce crystalline precipitates that are insoluble in water. The coordinated structures of the crystalline precipitates are shown in chemical structures 1 through 28. The crystalline precipitates produced thereby are metal chelate compounds. The

crystalline precipitated metal chelate compounds are made minute and one, two or more different metal chelate compounds are suspended or dispersed in water to generate aqueous lubricants.

In the structures, M represents divalent zinc, divalent or trivalent manganese, divalent or trivalent iron, trivalent, tetravalent or pentavalent molybdenum, divalent $[(\text{MoOS})_2]^{2+}$, divalent $[\text{Mo}_2\text{S}_4]^{2+}$, divalent or tetravalent tin, trivalent or pentavalent antimony, divalent MoO or monovalent MoOS.

In chemical structures 1-12, 15, 16, 18, 19 and 28, n varies depending on the M species. For example, when M is zinc, tin or antimony, n=1 or 2; when M is manganese or iron, n=1, 2 or 3; and when M is molybdenum, n=1 or 2. In the case of zinc, for example, if n=1, the chelate ligand having sulfur as a coordinating atom only coordinates with two of the four coordination sites of zinc. In this case, the two remaining coordination sites coordinate with a chelate ligand that does not have sulfur as a coordinating atom, such as a hydroxide ion, condensed phosphate, polycarboxylic high molecular activator and/or polyoxycarboxylic acid. Generally, this means that when the number of coordination sites of the metal M in the following chemical structures cannot be filled by a chelate ligand having sulfur as a coordinating atom, a chelate ligand containing no sulfur, such as hydroxide ion, condensed phosphate, polycarboxylic high molecular activator and/or polyoxycarboxylic acid, coordinates to the coordination sites

of the metal that are not coordinated with the chelate ligand having sulfur as a coordinating atom.

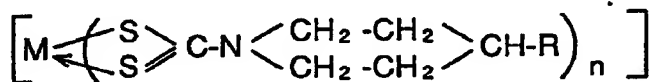
[Chemical formula 1]

5



wherein R_1 and R_2 may be the same group, in which case $R_1 (=R_2)$ is H, $-CH_3$, $-C_2H_5$, $-C_3H_7$ (straight chain), iso- C_3H_7 , $-C_4H_9$ (straight chain), iso- C_4H_9 , tert- C_4H_9 or $-C_6H_5$. R_1 and R_2 also may be different, and when R_1 is H, R_2 is $-CH_3$, $-C_2H_5$, $-C_3H_7$ (straight chain), iso- C_3H_7 , $-C_4H_9$ (straight chain), iso- C_4H_9 , tert- C_4H_9 or $-C_4H_9$ (straight chain). When R_1 is $-CH_3$ or $-C_2H_5$, R_2 is $-C_6H_5$.

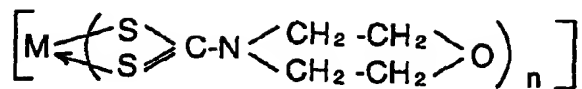
[Chemical formula 2]



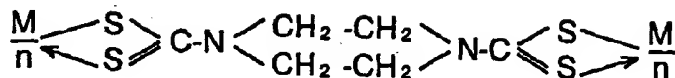
wherein R is H, $-CH_3$ or $-C_2H_5$.

20

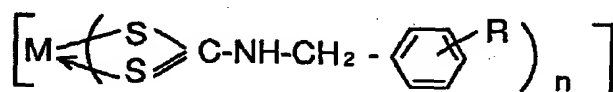
[Chemical formula 3]



[Chemical formula 4]

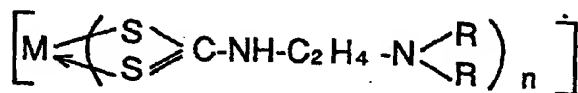


[Chemical formula 5]



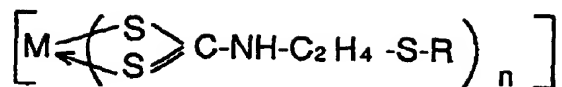
wherein R is ortho-NO₂, para-NO₂, meta-OCH₃, meta-CH₃ or meta-C₂H₅.

[Chemical formula 6]



wherein R is -CH₃, -C₂H₅, -C₃H₇ (straight chain) or iso-C₃H₇.

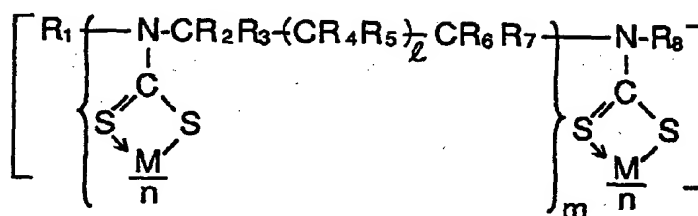
[Chemical formula 7]



wherein R is a hydrogen atom or an alkyl group of 1-12 carbon atoms.

5

[Chemical formula 8]



When $m=1$ and $l=0$:

$R_1\text{-R}_3$, and $R_6\text{-R}_8$ are H,

R_1 is $-\text{CH}_3$ and $R_2\text{-R}_3$, $R_6\text{-R}_8$ are H,

R_1 is $-\text{C}_2\text{H}_5$ and $R_2\text{-R}_3$, $R_6\text{-R}_8$ are H,

R_1 is $-\text{C}_3\text{H}_5$ (straight chain) or iso- C_3H_5 and $R_2\text{-R}_3$, $R_6\text{-R}_8$

are H,

R_1 is $-\text{C}_4\text{H}_9$ (straight chain), iso- C_4H_9 or tert- C_4H_9 , and

$R_2\text{-R}_3$, $R_6\text{-R}_8$ are H,

R_2 and R_3 are $-\text{CH}_3$ and R_1 , $R_6\text{-R}_8$ are H,

R_2 and R_6 are $-\text{CH}_3$ and R_1 , R_3 , $R_7\text{-R}_8$ are H,

R_2 , R_3 , R_6 and R_7 are $-\text{CH}_3$ and R_1 and R_8 are H,

R_1 and R_8 are $-\text{CH}_3$ and $R_2\text{-R}_3$, $R_6\text{-R}_7$ are H,

R_1 and R_8 are $-C_2H_5$ and R_2-R_3 , R_6-R_7 are H,

R_1 and R_8 are $-C_3H_5$ (straight chain) or iso- C_3H_5 and R_2-R_3 ,

R_6 , R_7 are H,

R_1 and R_8 are $-C_4H_9$ (straight chain), iso- C_4H_9 or tert-

5 C_4H_9 and R_2-R_3 , R_6-R_7 are H, or

R_2 and R_6 are $-C_6H_5$ and R_1 , R_3 , R_7-R_8 are H.

When $m=1$ and $l=1$,

R_1-R_6 are H, or

R_4 and R_5 are $-CH_3$ and R_1-R_3 , R_6-R_7 are H.

10 When $m=1$ and $l=2-7$,

R_1-R_8 are H.

When $m=2-9$ and $l=0$,

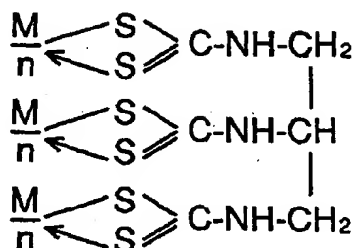
R_1-R_8 are H, or

R_2 is $-CH_3$ and R_1 , R_3 , R_6-R_8 are H.

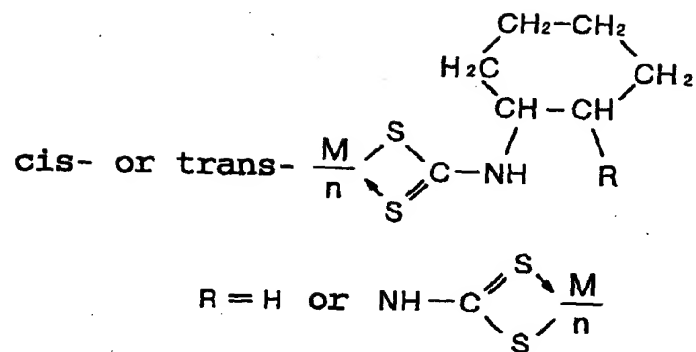
15 When $m=2-9$ and $l=1$,

R_1-R_8 are H.

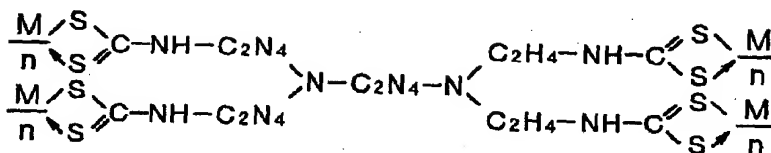
[Chemical formula 9]



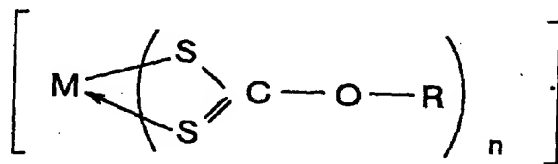
[Chemical formula 10]



[Chemical formula 11]

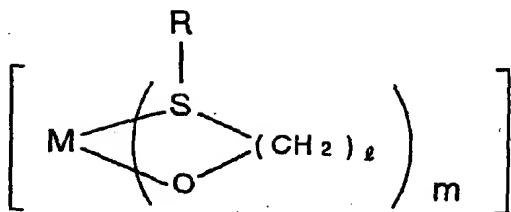


[Chemical formula 12]



wherein R is a linear or branched alkyl group of 1-12 carbon atoms.

[Chemical formula 13]

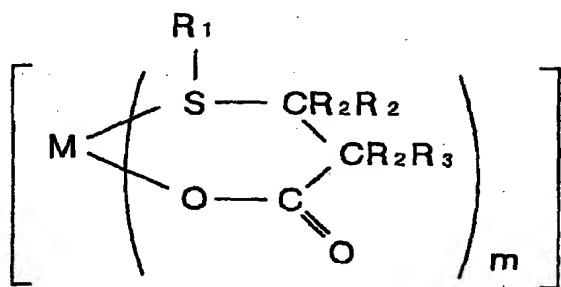


wherein R is H, l=2-3 and m=1, or

R is an alkyl group of 1-12 carbon atoms, l=2-3 and m=2-3.

5

[Chemical formula 14]



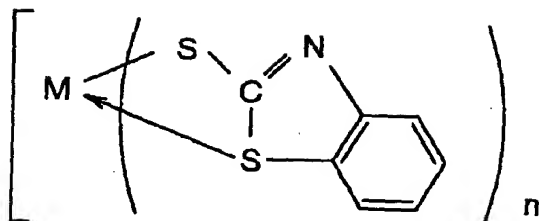
10 wherein R₁-R₃ are H and m=1,

R₁ is an alkyl group of 1-12 carbon atoms, R₂-R₃ are H and m=2,

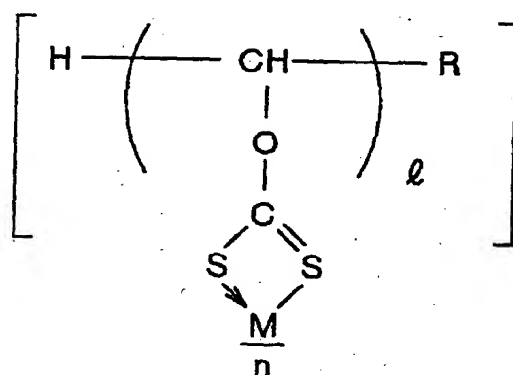
or

R₁-R₂ are H, R₃ is -NH₂ and m=2.

15 [Chemical formula 15]

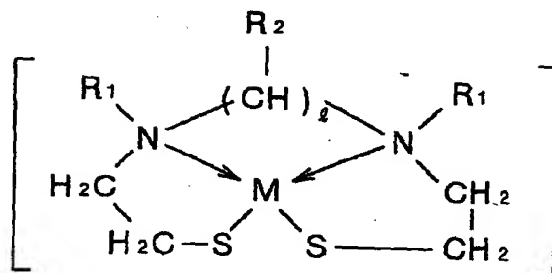


[Chemical formula 16]



wherein R is H, l=1-6 and n=2-3, or
R is -COOH, l=1-6 and n=2-3.

10 [Chemical formula 17]



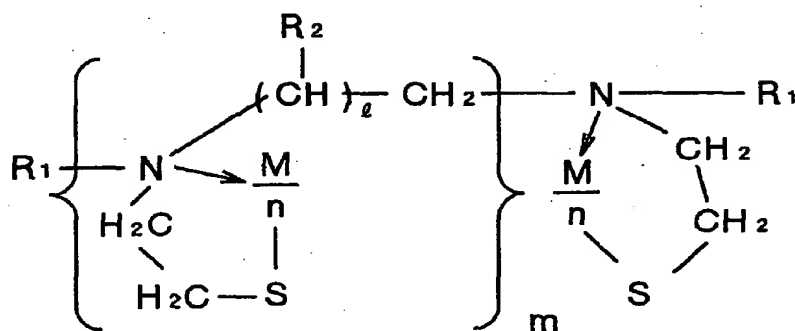
wherein R₁ and R₂ are H and l=2-12,

R_1 is H, R_2 is $-\text{CH}_3$ and $l=2-12$,

R_1 is $-\text{C}_2\text{H}_4\text{S}^-$, R_2 is H and $l=2-12$, or

R_1 is $-\text{C}_2\text{H}_4\text{S}^-$, R_2 is CH_3 and $l=2-12$.

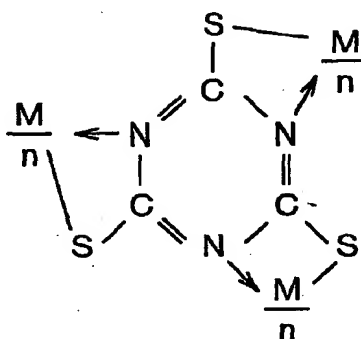
5 [Chemical formula 18]



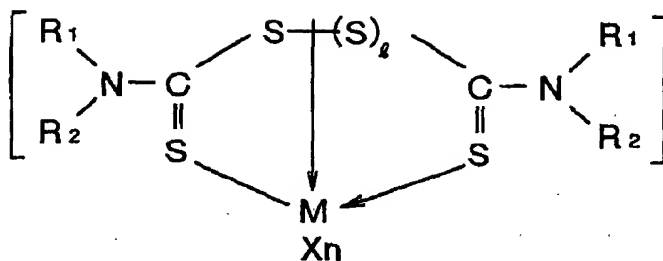
wherein R_1 and R_2 are H, $l=1-2$, $m=1-6$ and $n=2-3$, or

R_1 is $-\text{C}_2\text{H}_4\text{S}^-$, R_2 is H, $m=1-6$ and $n=1$.

[Chemical formula 19]



[Chemical formula 20]



wherein l is 0, 1, 2 or 3,

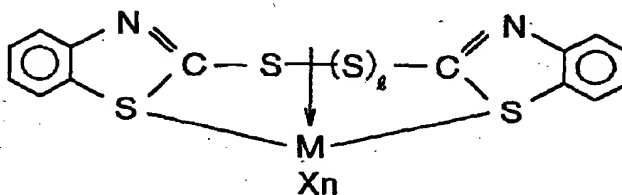
R_1 and R_2 are H, $-CH_3$, $-C_2H_5$, $-C_3H_7$ (straight chain) or $iso-C_3H_7$,

5 R_1 is H and R_2 is CH_3 , $-C_2H_5$, $-C_3H_7$ (straight chain) or $iso-C_3H_7$,
or

R_1 is $-C_2H_5$ and R_2 is $-C_6H_5$,

X is an anionic monodentate ligand and $n=2-3$.

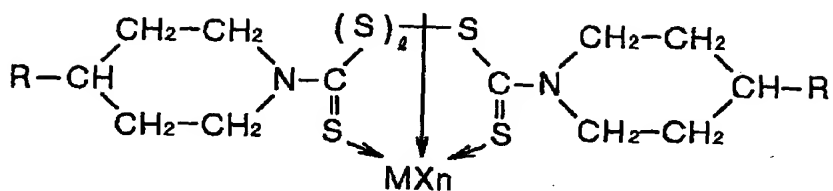
10 [Chemical formula 21]



wherein l is 0, 1, 2 or 3,

15 X is an anionic monodentate ligand and $n=2-3$.

[Chemical formula 22]

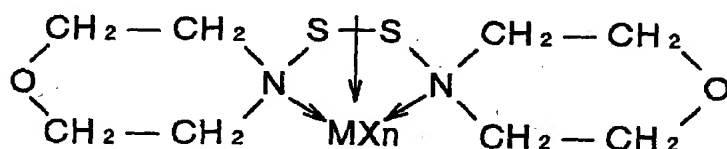


wherein R is H and l=1-4, or

R is an alkyl group of 1-3 carbon atoms and l=1-4,

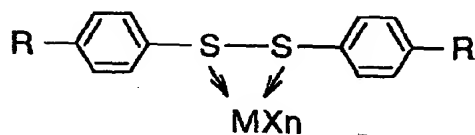
5 X is an anionic monodentate ligand and n=2-3.

[Chemical formula 23]



wherein X is an anionic monodentate ligand and n=2-3.

[Chemical formula 24]

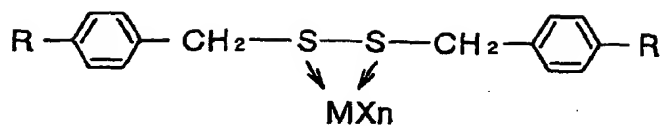


15

wherein R is H, -CH₃, -OCH₃, -OH or -C₆H₅,

X is an anionic monodentate ligand and n=2-3.

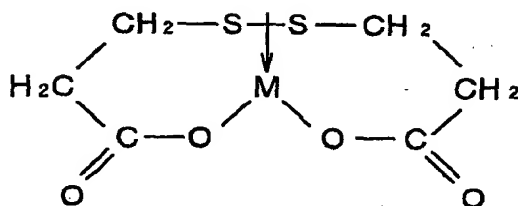
20 [Chemical formula 25]



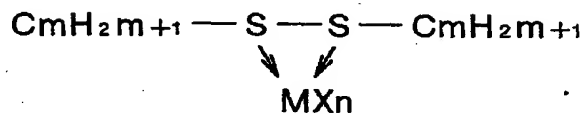
wherein R is H, -CH₃, -OCH₃ and -OH,

5 X is an anionic monodentate ligand and n=2-3.

[Chemical formula 26]

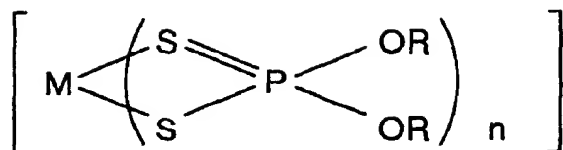


[Chemical formula 27]



15 wherein m=1-12, X is an anionic monodentate ligand and n=2-3.

[Chemical formula 28]



wherein R is an alkyl group (C=1-12).

5 As clearly seen in chemical structures 1-28, in these metal
chelate compounds, at least one coordinating atom is sulfur,
which is chelated to at least one metal species selected from
among zinc, manganese, iron, molybdenum, tin and antimony. The
metal chelate compounds are hydrophobic fine particles, but at
10 pH 8.0-13.0, they disperse in water with anionic or non-ionic
surfactants and are maintained as stable dispersions in water.
If the metal chelate compounds are instead made minute, they can
be suspended for relatively long periods without precipitation,
even if a surfactant or the like is not used, and a physical
15 suspension also can be created by stirring or agitation. Thus,
an aqueous lubricant can be realized that does not require any
oil or organic solvent.

When the suspension or dispersion is coated onto either
or both a metal material surface and the molding side of a metal
20 mold, a lubricating film is formed on the coated surface. The
lubricating film adheres well to the surface and does not easily
peel from the surface during plastic working of the metal material.
It also has satisfactory lubricating properties and effectively

prevents seizing of the material and the mold. This lubricant has the advantage of adhering well to the coated metal surface as long as no oil is present on that surface, and also has the feature of not requiring preparatory steps, such as degreasing and washing. In addition, the lubricant requires no special management and can be recycled, so that the only management necessary is re-supply of the consumed portion. Stringent washing is not required, even when electron beam welding is performed after plastic working.

A number of methods can be used for the application, and for example, application onto the surface of the material can be accomplished by any desired method, such as immersion of the material in the lubricant, painting with a brush or spraying. For application to the molding side of the mold, the method employed may be painting with a brush, spraying or the like.

The lubricant-coated material or mold may be permitted to stand so that the lubricant may naturally dry, but if necessary, it may be forcibly dried. The method employed for forcibly drying the lubricant may be any desired method, such as exposure to hot air, preheating the material or mold or drying by high-frequency heating. The extent of drying can be adjusted, if necessary, to achieve total dryness or partial dryness. The extent of drying can be adjusted, as desired, by varying the drying temperature and drying time.

The metal chelate compound may be produced in solution

instead of preparing the metal chelate compound beforehand and adding it to water. That is, the lubricant used can have one or more chelating agents in which at least one of the coordinating atoms is sulfur, and adding thereto a salt, oxide or hydroxide of one or more metals selected from among zinc, manganese, iron, molybdenum, tin and antimony and an anionic or non-ionic surfactant. This type of lubricant can be used in exactly the same manner.

In the case of the metal chelate compounds represented by chemical structures 1-28, chelate ligands, in which sulfur is a coordinating atom, may be coordinated to all the coordination sites of the metals. Alternatively, chelate ligands, in which sulfur is a coordinating atom, may be coordinated to only some of the coordination sites of the metals, and other ligands, in which sulfur is not a coordinating atom, are coordinated to the remaining coordination sites. Suitable examples of ligands in which sulfur is not a coordinating atom are hydroxide ions, condensed phosphate, polycarboxylic high molecular activators and/or polyoxycarboxylic acid. Chemical structures 1-28 also show compounds in which chelate ligands having sulfur as a coordinating atom, as represented in the structures, are coordinated to only some of the coordination sites of the metals.

When a phosphate film has been formed on the metal surface, and the metal is immersed in an aqueous solution of any of the chelate ligands (the compounds adjacent to M in the structures)

represented in chemical structures 1-28, the chelate ligand having sulfur as a coordinating atom coordinates with the zinc ion or iron ion present in the phosphate film to produce a crystalline multi-ligand metal chelate compound on the phosphate film, which results in a lubricating effect.

When a phosphate film has been formed on the metal surface, and the metal is immersed in an aqueous solution of a metal chelate compound in which a chelate ligand having sulfur as a coordinating atom is coordinated to only some of the coordination sites of the metal, while other chelate ligands are coordinated to the remaining coordination sites, a crystalline polynuclear chelate compound is produced on the phosphate film, resulting in a lubricating effect.

When the lubricating film includes two or more different types of metal chelate compounds, their lubricating effects are synergistic, so that a highly satisfactory effect is achieved.

The lubricant described above forms a lubricating film by strong adhesion to surfaces of primarily iron, especially steel and iron alloys, but it can also be used for non-ferrous metals, such as aluminum.

Various additives, such as pH adjusters, viscosity controllers, preservatives, antifoaming agents and the like may also be added to the lubricant. It is particularly preferred to add soluble condensed phosphate salts, fatty acid sodium salts, fatty acid potassium salts, soluble polycarboxylic high

molecular activators and/or soluble polyoxycarboxylic acid salts. These compounds increase the dispersability of the metal chelate compound in water, and improve the adhesive strength of the lubricating film to the metal surface.

5 Experimental examples will now be explained.

[Experimental Example 1]

150 g of zinc bis-(N,N-diethyldithiocarbamate) was added to a prepared solution of 20 g of sodium stearate in 100 ml of warm water and the mixture was gently stirred to produce an aqueous lubricant.

The following examples are analogous as aqueous lubricants to this Experimental Example 1. The metal chelate compound is not limited to zinc bis-(N,N-diethyldithiocarbamate), and it may be replaced with any of the species represented by chemical structures 1-28, such as N,N-dibutyldithiocarbamate oxymolybdenumsulfate.

Sodium stearate was used as the anionic or non-ionic surfactant in this experimental example, but adjustment to pH 8.0-13.0 with any other well-known anionic or non-ionic surfactants, such as sodium salts of fatty acids and/or potassium salts of fatty acids, can effect adequate dispersion of the metal chelate compound in water. A similar aqueous lubricant can be obtained by making the metal chelate compound fine, adding water and stirring it to create a suspension.

[Experimental Example 2]

A 50 g/200 ml aqueous solution of zinc sulfate heptahydrate was added to a 78 g/300 ml aqueous solution of sodium N,N-diethyldithiocarbamate trihydrate, while stirring, to prepare a suspension of zinc bis-(N,N-diethyldithiocarbamate). Separately, a suspension was created by dispersing 100 g of N,N-dibutyldithiocarbamate oxymolybdenumsulfate in a warm solution (500 ml) containing 20 g of sodium stearate, 20 g of sodium tripolyphosphate and 20 g of a polycarboxylic high molecular activator. Both suspensions were mixed together by stirring to obtain an aqueous lubricant.

The following examples are analogous as aqueous lubricants to this Experimental Example 2. In addition to the aqueous solution of sodium N,N-diethyldithiocarbamate trihydrate, the aqueous solution of zinc sulfate hydrate that produces a metal chelate compound may be replaced with another water-soluble zinc salt or zinc hydroxide compound. It may also be replaced with a water-soluble salt of manganese, iron, molybdenum, tin or antimony. Experimental Example 2 differs from Experimental Example 1 primarily in using a metal chelate compound of two or more different metals, and zinc and molybdenum are used here. Any combination of two or more metals from among zinc, manganese, iron, molybdenum, tin and antimony may be used. The chelate ligand used can be any of those represented in

chemical structures 1-28.

Sodium tripolyphosphate need not be included, but its addition will improve the dispersability of the metal chelate compound. The polycarboxylic activator also need not be included, but its addition will improve adhesion of the lubricating film to the metal. A soluble polyoxycarboxylic acid salt may be added instead of a polycarboxylic high molecular activator.

[Experimental Example 3]

A solution of 45.3 g of sodium N,N-diethyldithiocarbamate trihydrate and 8.5 g of sodium hydroxide dissolved in 200 ml of water was slowly poured into a 57.8 g/300 ml aqueous solution of zinc sulfate heptahydrate while stirring to prepare a suspension of zinc mono-(N,N-diethyldithiocarbamate)-hydroxo aqua.

Zinc mono-(N,N-diethyldithiocarbamate)-hydroxo aqua has a chelate ligand with sulfur as a coordinating atom strongly coordinated to some of the coordination sites of zinc, and sodium hydroxide ion weakly coordinated to the remaining coordination sites. Zinc mono-(N,N-diethyldithiocarbamate)-hydroxo aqua can be dispersed in water with a sodium fatty acid salt and/or potassium fatty acid salt, such as sodium stearate.

The following examples are analogous as aqueous lubricants to this Experimental Example 3. The chelate ligand

having sulfur as a coordinating atom that chelates to some of the coordination sites of the metal can be replaced with any desired ligand represented in chemical structures 1-28. The sodium hydroxide ion that coordinates to the remaining 5 coordination sites can be replaced with any other desired hydroxide ion, except for sulfur.

As in the above experimental examples, soluble condensed phosphate salts, soluble polycarboxylic high molecular activators and/or soluble polyoxycarboxylic acid salts may be added when necessary.

[Experimental Example 4]

18 g/100 ml of sodium tripolyphosphate was slowly poured into a 57.8 g/300 ml aqueous solution of zinc sulfate heptahydrate while stirring to prepare a suspension of a zinc tripolyphosphate crystalline precipitate. 45.3 g/200 ml of sodium N,N-diethyldithiocarbamate trihydrate was then slowly poured into the suspension, while stirring, to give a crystalline precipitate of zinc mono-(N,N-diethyldithiocarbamate)-triphosphate (hereunder, "G").

Separately, a suspension was prepared by dispersing 100 g of N,N-dibutyldithiocarbamate oxymolybdenumsulfate in a solution of 20 g of sodium stearate, 10 g of sodium tripolyphosphate and 12 g of a polycarboxylic activator dissolved 25 in 500 ml of hot water (hereunder, "H"). G and H were mixed

together by stirring to obtain a yellow dispersion that was used as a lubricant.

The zinc mono-(N,N-diethyldithiocarbamate)-triphosphate has a chelate ligand, with sulfur as a coordinating atom, strongly coordinated to some of the coordination sites of zinc, and sodium tripolyphosphate weakly coordinated to the remaining coordination sites via an oxygen anion. The species weakly coordinated to the remaining coordination sites via the oxygen anion is not limited to a condensed phosphate, such as sodium tripolyphosphate, and it may be replaced with a polycarboxylic high molecular activator and/or polyoxycarboxylic acid.

Experimental Example 4 differs from Experimental Example 3 primarily in using a metal chelate compound of two or more different metals, and zinc and molybdenum are used here. Alternatively, any combination of two or more metals selected from among zinc, manganese, iron, molybdenum, tin and antimony may be used. For two or more different metal chelate compounds, they may both have chelate ligands, in which sulfur is a coordinating atom, coordinated to some of the coordination sites of the metal, but as explained above, either of the metal chelate compounds may also have a chelate ligand with sulfur as a coordinating atom that coordinates to all the coordination sites of the metal. The chelate ligand used can be any of those represented in chemical structures 1-28.

[Experimental Example 5]

A metal material on which a phosphate film had already been formed was immersed in a warm solution of 5% sodium N,N-diethyldithiocarbamate (pH 10), and the crystalline multi-ligand zinc chelate compound produced on the phosphate film was used as a lubricant.

By adjusting the pH to a range of 6.5-13.5, the sodium N,N-diethyldithiocarbamate (ligand with sulfur as a coordinating atom) coordinates with the zinc ion or iron ion in the phosphate film to produce a crystalline multi-ligand zinc or iron chelate compound on the phosphate film, thus forming a lubricating film. The chelate ligands may be any desired ones represented by chemical structures 1-28.

[Experimental Example 6]

A mixed aqueous solution (150 ml) of 31 g of sodium N,N-diethyldithiocarbamate trihydrate and 5.9 g of sodium hydroxide was slowly poured into 40 g/200 ml of zinc sulfate heptahydrate, while stirring, to prepare a suspension of zinc mono-(N,N-diethyldithiocarbamate)-hydroxo aqua (hereunder, "I") (pH 11.5-12.0). A metal material, on which a phosphate film had already been formed, was immersed in I at 40-50°C and stirred for 30-60 seconds to produce a crystalline polynuclear zinc chelate on the phosphate film, which was used as the lubricating

film.

A crystalline polynuclear zinc chelate can also be disposed on the phosphate film by immersing the metal material with a phosphate film already formed thereon in any of the lubricants obtained in Experimental Examples 1 to 4. The chelate ligand can be any one represented in chemical structures 1-28, and it is particularly preferred for a chelate ligand with sulfur as a coordinating atom to be coordinated to some of the coordination sites of the metal and chelate ligands without sulfur coordinated to the remaining coordination sites. In such cases, the sulfur-containing chelate ligand coordinates with the metal while the non-sulfur-containing chelate ligand reacts with the zinc ions or iron ions in the phosphate film to produce a crystalline polynuclear metal chelate compound.

Each of the lubricants prepared in Examples 1-6 was coated onto the perforated side of a perforated testing billet (for area reduction of 12%: a cylindrical member was used and having an inner diameter of 15 mm, an outer diameter of 29.9 mm and a length of 50 mm), and dried by exposure to 150°C hot air for 60 seconds. The time required for this treatment was about 2 minutes.

For comparison, a phosphate film was formed on the same type of billet and a metallic soap film was disposed on top of the phosphate film (Comparative Example 1). The time required for this treatment was over 30 minutes.

For further comparison, a working oil (with lubricating

adjuvant added) was coated onto the same type of billet (Comparative Example 2). The time required for this treatment was about 30 seconds.

Each of the billets was subjected to a ball-push test.

5 This test measured the load required for plastic deformation of a billet when an iron ball with a larger diameter than the diameter of the perforation in the billet was forcibly pushed through the billet perforation. The lubricating performance was evaluated based upon the surface condition of the inner diameter of the billet. A smaller load indicates more satisfactory lubrication, allowing smoother plastic deformation. The results (maximum loads) are listed in the following table. In the table, the area reductions are the rates of change in the billet perforations before and after deformation, with larger values indicating a higher degree of deformation, i.e. heavy working. It will be appreciated that better lubrication is achieved when the surface condition has no seizing or, where seizing occurs, the length is shorter.

20 Values in table indicate loads (units: t)

Area Reduction	6%	8%	10%	12%	Treatment Time
Exp. Example 1	19.625	NG	NG	NG	2 min.
Exp. Example 2	11.676	15.858	20.721	NG	2 min.
Exp. Example 3	13.930	16.710	NG	NG	2 min.
Exp. Example 4	9.265	9.876	11.934	15.992	2 min.
Exp. Example 5	10.176	13.945	17.116	20.174	2 min.
Exp. Example 6	8.096	9.371	10.767	15.130	2 min.
Comp. Example 1	8.383	10.128	14.068	18.060	30 min.
Comp. Example 2	21.391	NG	NG	NG	30 sec.

In the table, NG indicates seizing between the iron ball and the billet, showing that a satisfactory surface condition was not obtained. Not all of the experimental examples of the invention were suitable for heavy working, but light working was possible in all of the experimental examples, and satisfactory lubricating performance was confirmed in all of the experimental examples, as compared with using oil. Heavy working is also possible by appropriate selection among the lubricants of the invention.

These examples confirm that results comparable to Comparative Example 1 can be obtained with adequate functionality by simple coating and formation of a lubricating film. Because the treatment times were only about 2 minutes, treatment is possible within a very short time. The performance difference was particularly notable in comparison with Comparative Example 2, thus confirming that working is possible by these examples even in cases in which working was not possible by Comparative Example 2.

Effect of the Invention

With the lubricating solution of the invention, effective lubricating films can be formed by a simple application process, which films exhibit performance comparable to that achieved by the troublesome and difficult procedure of forming phosphate

films and disposing metallic soap films on the metal object. Because the present invention does not use oil, the problems of working environment contamination and the need for subsequent degreasing treatment do not occur. Because a simple application
5 is sufficient, a solution is also provided to the problems of large equipment requirements and the need to perform simultaneous treatment of large volumes of materials. Apparatus can therefore be installed for formation of lubricating films in the narrow spaces adjacent to the apparatus for plastic working of materials, forming the lubricating films in sequence with the plastic working apparatus cycles without requiring extra storage between the two treatments, so that it becomes possible to shorten lead times.

10
T05TF T048866